Characterization of Covalently Functionalized Single-Walled Carbon Nanotubes*

Liling Zeng
Andrew R. Barron

This work is produced by OpenStax-CNX and licensed under the
Creative Commons Attribution License 3.0†

Abstract

A summary of common methods used to characterize chemically functionalized single-walled carbon nanotubes (SWNTs).

1 Introduction

Characterization of nanoparticles in general, and carbon nanotubes in particular, remains a technical challenge even though the chemistry of covalent functionalization has been studied for more than a decade. It has been noted by several researchers that the characterization of products represents a constant problem in nanotube chemistry. A systematic tool or suites of tools are needed for adequate characterization of chemically functionalized single-walled carbon nanotubes (SWNTs), and is necessary for declaration of success or failure in functionalization trials.

So far, a wide range of techniques have been applied to characterize functionalized SWNTs: infra red (IR), Raman, and UV/visible spectroscopies, thermogravimetric analysis (TGA), atomic force microscopy (AFM), transmission electron microscopy (TEM), X-ray photoelectron spectroscopy (XPS), etc. A summary of the attribute of each of the characterization method is given in Table 1.

Method	Sample	Information	Limitations
TGA	solid	functionalization ratio	no evidence for covalent functionalization, not specific
continued on next page			

*Version 1.4: Jan 7, 2010 1:55 pm -0600
†http://creativecommons.org/licenses/by/3.0/

XPS	solid	elements, functionalization ratio	no evidence of covalent functionalization, not specific, quantification complicated
Raman	solid	sp^3 indicated by D mode	not specific, quantification not reliable
Infra red (IR)	solid for (ATR-IR) or solution	substituent groups	no direct evidence for covalent functionalization, quantification not possible
UV/visible	solution	sidewall functionalization	not specific or quantitative, need highly dispersed sample
Solution NMR	solution	substituents	no evidence of covalent functionalization, high solubility of sample
Solid state NMR	solid	substituents, sp^3 molecular motions, quantification at high level of funcitionalization	high functionalization needed, long time for signal acquisition, quantification not available for samples with protons on side chains
AFM	solid on substrate	topography	only a small portion of sample characterized, no evidence of covalent functionalization, no chemical identity
TEM	solid on substrate	image of sample distribution dispersion	only a small portion of sample characterized, no evidence of covalent functionalization, no chemical identity dispersion information complicated

continued on next page

STM	solid on substrate	distribution	no chemical identity of functional groups small portion of sample conductive sample only

Table 1: Common characterization methodology for functionalized SWNTs.

2 Elemental and Physical Analysis

2.1 Thermogravimetric analysis (TGA)

Thermogravimetric analysis (TGA) is the mostly widely used method to determine the level of sidewall functionalization. Since most functional groups are labile or decompose upon heating, while the SWNTs are stable up to 1200 °C under Ar atmosphere. The weight loss at 800 °C under Ar is often used to determine functionalization ratio using this indirect method. Unfortunately, quantification can be complicated with presence of multiple functional groups. Also, TGA does not provide direct evidence for covalent functionalization since it cannot differentiate between covalent attachment and physical adsorption.

2.2 X-ray photoelectron spectroscopy (XPS)

XPS confirms the presence of different elements in functionalized SWNTs. This is useful for identification of heteroatom elements such as F and N, and then XPS can be used for quantification with simple substituent groups and used indirectly. Deconvolution of XPS is useful to study fine structures on SWNTs. However, the overlapping of binding energies in the spectrum complicates quantification.

3 Spectroscopy

3.1 Raman spectroscopy

Raman spectroscopy is very informative and important for characterizing functionalized SWNTs. The tangential G mode (*ca.* $1550 - 1600$ cm^{-1}) is characteristic of sp^2 carbons on the hexagonal graphene network. The D-band, so-called disorder mode (found at *ca.* 1295 cm^{-1}) appears due to disruption of the hexagonal sp^2 network of SWNTs. The D-band was largely used to characterize functionalized SWNTs and ensure functionalization is covalent and occurred at the sidewalls. However, the observation of D band in Raman can also be related to presence of defects such as vacancies, 5-7 pairs, or dopants. Thus, using Raman to provide evidence of covalent functionalization needs to be done with caution. In particular, the use of Raman spectroscopy for a determination of the degree of functionalization is not reliable.

It has been shown that quantification with Raman is complicated by the distribution of functional groups on the sidewall of SWNTs. For example, if fluorinated-SWNTs (F-SWNTs) are functionalized with thiol or thiophene terminated moieties, TGA shows that they have similar level of functionalization. However, their relative intensities of D:G in Raman spectrum are quite different. The use of sulfur substituents allow for gold nanoparticles with 5 nm in diameter to be attached as a "chemical marker" for direct imaging of the distribution of functional groups. AFM and STM suggest that the functional groups of thio-SWNTs are group together while the thiophene groups are widely distributed on the sidewall of SWNTs. Thus the difference is not due to significant difference in substituent concentration but on substituent distribution, while Raman shows different D:G ratio.

3.2 Infra red spectroscopy

IR spectroscopy is useful in characterizing functional groups bound to SWNTs. A variety of organic functional groups on sidewall of SWNTs have been identified by IR, such as COOH(R), -CH_2, -CH_3, -NH_2, -OH, etc. However, it is difficult to get direct functionalization information from IR spectroscopy. The C-F group

has been identified by IR in F-SWNTs. However, C-C, C-N, C-O groups associated with the side-wall functionalization have not been observed in the appropriately functionalized SWNTs.

3.3 UV/visible spectroscopy

UV/visible spectroscopy is maybe the most accessible technique that provides information about the electronic states of SWNTs, and hence functionalization. The absorption spectrum shows bands at *ca.* 1400 nm and 1800 nm for pristine SWNTs. A complete loss of such structure is observed after chemical alteration of SWNTs sidewalls. However, such information is not quantitative and also does not show what type of functional moiety is on the sidewall of SWNTs.

3.4 Nuclear magnetic resonance

NMR can be considered as a "new" characterization technique as far as SWNTs are concerned. Solution state NMR is limited for SWNT characterization because low solubility and slow tumbling of the SWNTs results in broad spectra. Despite this issue, there are still solution ^{1}H NMR reported of SWNTs functionalized by carbenes, nitrenes and azomethine ylides because of the high solubility of derivatized SWNTs. However, proof of covalent functionalization cannot be obtained from the ^{1}H NMR. As an alternative, solid state ^{13}C NMR has been employed to characterize several functionalized SWNTs and show successful observation of sidewall organic functional groups, such as carboxylic and alkyl groups. But there has been a lack of direct evidence of sp^{3} carbons on the sidewall of SWNTs that provides information of covalent functionalization.

Solid state ^{13}C NMR has been successfully employed in the characterization of F-SWNTs through the direct observation of the sp^{3}*C*-F carbons on sidewall of SWNTs. This methodology has been transferred to more complicated systems; however, it has been found that longer side chain length increases the ease to observe sp^{3}*C*-X sidewall carbons.

Solid state NMR is a potentially powerful technique for characterizing functionalized SWNTs because molecular dynamic information can also be obtained. Observation that higher side chain mobility can be achieved by using a longer side chain length offers a method of exploring functional group conformation. In fact, there have been reports using solid state NMR to study molecular mobility of functionalized multi-walled carbon nanotubes.

4 Microscopy

AFM, TEM and STM are useful imaging techniques to characterize functionalized SWNTs. As techniques, they are routinely used to provide an "image" of an individual nanoparticle, as opposed to an average of all the particles.

4.1 Atomic force microscopy

AFM shows morphology on the surface of SWNTs. The height profile on AFM is often used to show presence of functional groups on sidewall of SWNTs. Individual SWNTs can be probed by AFM and sometimes provide information of dispersion and exfoliation of bundles. Measurement of heights along an individual SWNT can be correlated with the substituent group, i.e., the larger an alkyl chain of a sidewall substituent the greater the height measured. AFM does not distinguish whether those functional groups are covalently attached or physically adsorbed on the surface of SWNTs.

4.2 Transmission electron microscopy

TEM can be used to directly image SWNTs and at high resolution clearly shows the sidewall of individual SWNT. However, the resolution of TEM is not sufficient to directly observe covalent attachment of chemical modification moieties, i.e., to differentiate between sp^{2} and sp^{3} carbon atoms. TEM can be used to provide information of functionalization effect on dispersion and exfoliation of ropes.

Samples are usually prepared from very dilute concentration of SWNTs. Sample needs to be very homogeneous to get reliable data. As with AFM, TEM only shows a very small portion of sample, using them to characterize functionalized SWNTs and evaluate dispersion of samples in solvents needs to be done with caution.

4.3 Scanning tunneling microscopy

STM offers a lot of insight on structure and surface of functionalized SWNTs. STM measures electronic structure, while sometimes the topographical information can be indirectly inferred by STM images. STM has been used to characterize F-SWNTs gold-marked SWNTs, and organic functionalized SWNTs. Distribution of functional groups can be inferred from STM images since the location of a substituent alters the localized electronic structure of the tube. STM images the position/location of chemical changes to the SWNT structure. The band-like structure of F-SWNTs was first disclosed by STM.

STM has the same problem that is inherent with AFM and TEM, that when using small sample size, the result may not be statistically relevant. Also, chemical identity of the features on SWNTs cannot be determined by STM; rather, they have to be identified by spectroscopic methods such as IR or NMR. A difficulty with STM imaging is that the sample has to be conductive, thus deposition of the SWNT onto a gold (or similar) surface is necessary.

5 Bibliography

- L. B. Alemany, L. Zhang, L. Zeng, C. L. Edwards, and A. R. Barron, *Chem. Mater.*, 2007, **19**, 735.
- J. L. Bahr and J. M. Tour, *J. Mater. Chem.*, 2002, **12**, 1952.
- M. S. Dresselhaus, G. Dresselhaus, and A. Jorio, *J. Phys. Chem. C*, 2007, **111**, 17887.
- A. Hirsch, *Angew. Chem. Int. Ed.*, 2002, **41**, 1853.
- M. Holzinger, J. Abraham, P. Whelan, R. Graupner, L. Ley, F. Hennrich, M. Kappes, and A. Hirsch, *J. Am. Chem. Soc.*, 2003, **125**, 8566.
- K. F. Kelly, I. W. Chiang, E. T. Mickelson, R. H. Hauge, J. L. Margrave, X. Wang, G. E. Scuseria, C. Radloff, and N. J. Halas, *Chem. Phys. Lett.*, 1999, **313**, 455.
- V. N. Khabashesku, W. E. Billups, and J. L. Margrave, *Acc. Chem. Res.*, 2002, **35**, 1087.
- F. Liang, L. B. Alemany, J. M. Beach, and W. E. Billups, *J. Am. Chem. Soc.*, 2005, **127**, 13941.
- D. Tasis, N. Tagmatarchis, A. Bianco, and M. Prato, *Chem. Rev.*, 2006, **106**, 1105.
- H-L. Wu, Y-T. Yang, C-C. M. Ma, and H-C. Kuan, *J. Polym. Sci. A. Polym. Chem.*, 2005, 6084.
- L. Zeng, L. Zhang, and A. R. Barron, *Nano Lett.*, 2005, **5**, 2001.
- L. Zhang, J. Zhang, N. Schmandt, J. Cratty, V. N. Khabashesku, K. F. Kelly, and A. R. Barron, *Chem. Commun.*, 2005, 5429.